

High-pressure synthesis of boron-rich chalcogenides B₁₂S and B₁₂Se

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ABSTRACT

Two boron-rich chalcogenides B₁₂S and B₁₂Se isostructural to α -rhombohedral boron were synthesized by chemical reaction of the elements at high-pressure – high-temperature conditions. The crystal structures and stoichiometries of both compounds were confirmed by Rietveld refinement and elemental analysis. The experimental Raman spectra of B₁₂S and B₁₂Se were investigated for the first time. All observed Raman bands have been attributed to the theoretically calculated phonon modes, and the mode assignment has been performed.

INTRODUCTION

Boron-rich compounds isostructural to the α -rhombohedral boron (α -B₁₂) have become the subject of extensive theoretical¹⁻¹¹ and experimental¹²⁻²³ studies due to their unusual properties and potential technical applications.^{14,18,24} One of the most attractive features of boron-rich compounds is their outstanding mechanical properties. For instance, the reported hardness values of boron suboxide (B₁₂O₂) varies from 24 to 45 GPa.²⁵⁻³¹ Thus, boron suboxide is believed to be the hardest known oxide. Numerous works have been devoted to the investigation of its phase stability,^{3,32} compressibility,³²⁻³⁴ phonon³⁵⁻³⁷ and thermal^{34,38} properties, etc. However, other boron-rich chalcogenides, unlike boron suboxide, are still poorly studied.

Up to now, there have been only a few reports on the synthesis of boron subsulfide (B₁₂S_x)³⁹⁻⁴¹ and boron subselenide (B₁₂Se_x).⁴² In all these studies boron chalcogenides were synthesized by chemical reaction between elemental boron and sulfur/selenium in graphite or tantalum crucibles at high temperatures (1200-1600°C) in Ar atmosphere. The chalcogen content “x” varies from 0.9 to 1.3 in

B₁₂S_x, and from 0.9 to 1.1 in B₁₂Se_x. This is not surprising taking into account relatively low boiling temperatures of elemental sulfur (718 K) and selenium (958 K). Thus, the synthesis conditions (e.g. starting reagents ratios, maximum temperature, heating time, etc.) influence significantly the final stoichiometry of the products.

The loss of elemental sulfur and selenium during the high-temperature synthesis of boron chalcogenides can be prevented by applying high pressure. For instance, the high-pressure – high-temperature (HP-HT) synthesis provided a reproducible stoichiometry of many boron-rich compounds, such as B₁₂O₂,²¹ B₁₃N₂,¹⁷ B₅₀N₂,⁴³ and B₁₂As₂.⁴⁴ Thus, there are strong grounds to believe that HP-HT synthesis can ensure the reproducible chemical composition of boron-rich chalcogenides as well.

In the present work, we performed HP-HT synthesis of B₁₂S and B₁₂Se. The crystal structures and chemical compositions of both boron-rich chalcogenides were confirmed by Rietveld refinement of the powder X-ray diffraction data and elemental analysis. The Raman spectra of both compounds have been experimentally observed for the first time. Based on *ab initio* calculations, we assigned all observed Raman bands.

EXPERIMENTAL

Polycrystalline samples of B₁₂S and B₁₂Se were synthesized at 6 GPa and 2500 K by reaction of elemental boron (Grade I ABCR) with sulfur/selenium (both Alfa Aesar, 99.5%) powders mixed in 15:1/18:1 molar ratios. The compacted mixtures of starting reagents were loaded in hBN capsules (to isolate the reaction mixture from graphite heater) and placed in the high-temperature assembly of a toroid-type high-pressure apparatus.⁴⁵ After reaching the required pressure, the reaction mixtures were heated at 2500 K for 3 min, then gradually cooled down to 1500 K during 10 min and quenched. In order to remove unreacted boron, the recovered samples were treated with 3N nitric acid (ACS, Alfa Aesar) for 20 min at 370 K, washed by deionized water and dried at 400 K.

The structure of as-synthesized compounds was studied by angle-dispersive powder X-ray diffraction at Swiss-Norwegian Beamline BM01, ESRF.⁴⁶ The wavelength of the monochromatic beam from a bending magnet was set to 0.6866 Å. Diffraction patterns were acquired during the 20 s in Debye-Scherrer geometry with rotating quartz-glass capillary using PILATUS 2M detector (high purity LaB₆ was used as an internal standard). The crystal structures of both compounds were refined using Maud software⁴⁷ (Fig. 1).

The chemical composition of B₁₂S and B₁₂Se powders was studied by energy-dispersive X-ray analysis using scanning electron microscopes FEI Quanta 200F with X-Max EDS system (Oxford Instruments) and Hitachi SU8010 with SDD X-Maxⁿ EDX system (Oxford Instruments).

Raman spectra were measured in the 100-1600 cm⁻¹ range at ambient conditions using Horiba Jobin Yvon HR800 spectrometer calibrated with single-crystal cubic Si at room temperature. In order to check for possible resonant effects and/or photoluminescence, the measurements were performed at

two different excitation wavelengths: 473 nm and 633 nm. None of those phenomena have been observed for both compounds.

COMPUTATIONAL

The lattice parameters and atomic fractional coordinates of trigonal *R-3m* B₁₂S and B₁₂Se phases were optimized using a linear combination of atomic orbitals (LCAO) method implemented in CRYSTAL17 code.⁴⁸ The starting unit cell parameters and atom coordinates were taken from the experimental data. Improved all-electron double-zeta valence basis sets augmented by one set of polarization functions (pob-DZVP-rev2)⁴⁹ were chosen for boron, sulfur and selenium atoms. The accuracy of the calculation of the bielectronic Coulomb and exchange series was controlled by the set of tolerances, which were set to 10⁻⁸, 10⁻⁸, 10⁻⁸, 10⁻⁸, and 10⁻¹⁶. The Monkhorst-Pack scheme⁵⁰ for an 8×8×8 k-point mesh was used to integrate the Brillouin zone. Self-consistent field calculations were performed for hybrid DFT/HF WCGGA-PBE-16% functional.⁵¹ The percentage of 16% defines the Hartree-Fock admixture in the exchange part of DFT functional. The tolerance for the total energy change was set to 10⁻¹⁰.

To calculate B₁₂S and B₁₂Se phonon frequencies the direct (frozen-phonon) method implemented in CRYSTAL17 code^{53,54} was used. Raman intensities were calculated by using a coupled-perturbed Hartree-Fock/Kohn-Sham approach.^{54,55} The parameters of the optimized unit cell as well as atomic coordinates are collected in Table IV. Raman spectra were constructed by using the transverse optical (TO) modes and by adopting a pseudo-Voigt functional form⁵³ with a full width half maximum parameter set to 1.

RESULTS AND DISCUSSION

Crystal structures of B₁₂S and B₁₂Se

As typical α -B₁₂-related compounds, boron subsulfide and boron subselenide have rhombohedral symmetry (*R-3m* space group). The refined lattice parameters of B₁₂S and B₁₂Se were found to be in good agreement with the literature data (Table I). The unit cells of both compounds contain two independent boron atoms (in *18h* Wyckoff positions) and one sulfur/selenium atom (in *6c* Wyckoff position). All atomic coordinates and bond lengths are presented in Table II. The slightly distorted B₁₂-icosahedra are placed in the corners of the unit cell and on one of the unit cell main diagonals. S/Se atoms have a tetrahedral environment, including three B atoms belonging to three different boron icosahedra and one S/Se atom. According to the Rietveld refinement, the occupancies of *6c* sites by S and Se atoms were found to be 55% and 52%, respectively. Since all boron atoms constitute B₁₂-icosahedra, the occupancies of *18h* sites were fixed to 1.0 by default. The results of Rietveld refinement are in good agreement with energy-dispersive X-ray spectroscopy data: the

elemental compositions of B₁₂S_x and B₁₂Se_x are 92.5(1) at% B and 7.5(1) at% S and 92.6(1) at% B and 7.4(1) at% Se, i.e. $x \approx 1$ for both compounds. The final reliability factors (R_{wp}) were converged to 2.4 for B₁₂S and 0.5 for B₁₂Se, which indicates the excellent refinement level (Fig. 1).

The occupancies of $6c$ sites by S/Se atoms in B₁₂S and B₁₂Se refined in our study are close to the values obtained previously⁴⁰⁻⁴² (Table I). An attempt to place a single S/Se atom in $3b$ site (between two $6c$ sites) as it was proposed by Matkovich³⁹ resulted in not satisfactory profile fit convergence of the Rietveld refinement. Moreover, a replacement of the S/Se atom by B atom in $6c$ positions as it was done previously (e.g. B₁₂Se_{2-x}B_x)⁴² led to significant refinement deteriorations and, hence, high R_{wp} values. Thus, despite the partial occupancies of $6c$ sites in B₁₂S/B₁₂Se unit cells, we exclude a possibility of the S/Se atoms “slide” to $3b$ Wyckoff position or their partial replacement by boron atoms.

The ~50% occupancy of $6c$ site in B₁₂Se can be explained by the short Se-Se distance (2.02 Å) compared with the covalent radius of selenium (1.20 Å). In other words, B₁₂Se unit cell is too small to accommodate two Se atoms. However, the same logic cannot be applied for B₁₂S: S-S interatomic distance is 2.23 Å, whereas the covalent sulfur radius is 1.05 Å. The nature of the partial occupancies of $6c$ sites in B₁₂S remains unclear. As follows from Table I and Figure 2, the larger values of $6c$ site occupancy in boron subsulfides expectedly lead to the expansion of the lattice parameters (particularly, along the c -axis) and increase of the calculated densities. Unlike boron subsulfides, larger values of $6c$ site occupancy in boron subselenide surprisingly result in the lower calculated density and shrinking of the unit cell along the c -axis (Table I).

Raman spectroscopy study of B₁₂S and B₁₂Se

Unlike recently synthesized orthorhombic boron-rich chalcogenides, o -B₆S and o -B₆Se,⁵⁶ the Raman spectra of rhombohedral B₁₂S and B₁₂Se contain broad "noisy" bands in the 150-1250 cm⁻¹ range (Figure 3). According to the symmetry analysis, the acoustic and optic modes of boron-rich chalcogenides (42 atoms per unit cell) at Γ point can be presented as follows:

$$\Gamma_{\text{acoustic}} = A_{2u} + E_u$$

$$\Gamma_{\text{optic}} = 5A_{1g} + 5A_{2u} + 7E_u + 7E_g$$

$5A_{2u} + 7E_u$ are IR active modes; $5A_{1g} + 7E_g$ are Raman active modes for the D_{3d} point group.

The theoretically predicted Raman spectra of B₁₂S and B₁₂Se (at $T = 0$ K) are presented in Figure 3 (the positions of the theoretical Raman bands peaks are denoted by red dashed lines). The wavenumbers of the calculated Raman active phonon modes (ω^t) and the experimentally observed Raman bands (ω_0) are collected in Table III. As it follows from Fig. 3 and Table III, the theoretical and experimental data are in good agreement (as it was also observed in our previous Raman studies^{57,58}): the average difference between experimental and theoretical data is less than 3.5% for

B₁₂S (with a maximum value of 6.1% for 486.2 cm⁻¹ mode) and 1.5% for B₁₂Se (with a maximum value of 6.3% for 477.9 cm⁻¹ mode).

Using the visualization procedure built in MOLDRAW software,⁵⁹ the theoretically predicted phonon modes were attributed to the oscillations of two main structural elements of B₁₂X, where X = S/Se: triangle composed of 'B1' polar boron atoms (further, (B1)₃ unit) and X atom linked with three different 'B2' equatorial boron atoms (further, X-(B2)₃ unit) of B₁₂-icosahedra (Fig. 4).

The low-frequency A_{1g} and E_g modes (323.3 cm⁻¹ and 355.2 cm⁻¹ for B₁₂S and 199.6 cm⁻¹ and 231.5 cm⁻¹ for B₁₂Se) were referred to S/Se atoms oscillations (parallel and perpendicular to the *c*-axis) and the corresponding slight distortions in B₁₂-icosahedra. As it follows from Fig. 3b, A_{1g} and E_g modes have not been experimentally observed for B₁₂Se in the 150-400 cm⁻¹ range. The most probable explanation for this phenomenon may be the low intensity of the corresponding Raman bands and the low signal-to-noise ratio of the whole Raman spectrum.

The E_g modes of B₁₂S (458.2 cm⁻¹; 731.8 cm⁻¹) and B₁₂Se (449.6 cm⁻¹; 697.5 cm⁻¹) can be assigned to the tilting of (B1)₃ units and twisting oscillations of boron atoms in X-(B2)₃ unit leading to the tilting of the whole icosahedra around different unit cell directions. Meanwhile, the E_g modes of B₁₂S (754.1 cm⁻¹; 824.8 cm⁻¹; 1047.6 cm⁻¹; 1095.1 cm⁻¹) and B₁₂Se (711.7 cm⁻¹; 804.3 cm⁻¹; 1020.7 cm⁻¹; 1062.8 cm⁻¹) correspond to symmetric and asymmetric stretching/twisting oscillations of B–B bonds and stretching/rocking oscillations of B–X bonds in (B1)₃ and X-(B2)₃ units, respectively. It might be roughly assumed that vibrations of (B1)₃ units lead basically to the distortions of the icosahedra and, thus, to distortions of the intra-icosahedral bonds, whereas X-(B2)₃ vibrations result in vibration of the inter-icosahedral bonds. It should also be noted that the amplitude of the X-(B2)₃ vibrations are more significant at the highest frequencies, while for the (B1)₃ unit it is the opposite.

The A_{1g} modes of B₁₂S (625.0 cm⁻¹; 761.9 cm⁻¹) and B₁₂Se (645.3 cm⁻¹; 741.1 cm⁻¹) correspond to the symmetric "parasol" oscillations of X-(B2)₃ units, thus, resulting in distortions of the intra-icosahedral bonds. The A_{1g} modes of B₁₂S (1039.2 cm⁻¹) and B₁₂Se (1024.1 cm⁻¹) refer to the symmetric stretching of B–X bonds in X-(B2)₃ unit. Finally, the A_{1g} modes of B₁₂S (1093.7 cm⁻¹) and B₁₂Se (1038.4 cm⁻¹) were assigned to the symmetric stretching of (B1)₃ units. The two latter A_{1g} modes of both boron-rich chalcogenides lead to the distortions of the inter-icosahedral bonds.

Taking into account a good agreement between theoretical and experimental data the phonon modes assignment provided above for the calculated Raman spectra of both compounds is also fear-relevant for the experimental spectra. The accuracy of the present mode assignment suffers from the noise and broad Raman bands that lead to the bands overlapping. Nevertheless, these features are rather typical for the Raman spectra of other boron-rich compounds (e.g. B₄C⁶⁰ or B₁₂O₂³⁵). The more precise description of the phonon modes of B₁₂S and B₁₂Se might be done with the help of first- and second-order Raman scattering.³⁷

CONCLUSIONS

In the present work two new stoichiometric boron-rich chalcogenides, B₁₂S and B₁₂Se, were synthesized at high pressure – high temperature conditions and studied by powder X-ray diffraction and Raman spectroscopy at ambient pressure. According to the Rietveld refinement of synchrotron X-ray diffraction data, both B₁₂S and B₁₂Se have rhombohedral symmetry and belong to the space group *R-3m* (No 166). The observed Raman bands were assigned to the phonon modes and associated with the corresponding atomic movements.

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NOTES

The authors declare no competing financial interest.

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Table I. Lattice parameters, S(Se) occupation of *6c* sites and X-ray densities of B₁₂S and B₁₂Se : present work (*pw*) and literature data.

Compound	<i>a</i> , Å	<i>c</i> , Å	Occup., %	ρ, g/cm ³	Reference
B ₁₂ S	5.8196(2)	11.9653(5)	55	2.34	<i>pw</i>
	5.80	11.90	50*	2.33	[39]
	5.810(2)	11.94(2)	48.5	—	[40]
	5.8624(9)	12.147(4)	65	—	
	5.8379(6)	12.036(1)	59.9	2.36	[41]
	5.8307(5)	12.028(2)	60.9	2.37	
	5.8273(3)	12.025(2)	62	2.39	
B ₁₂ Se	5.9385(1)	11.9144(2)	52	2.90	<i>pw</i>
	5.9041(4)	11.947(1)	46.9**	2.97	[42]

* S atom occupies *3b* position (occupation is 100%), thus stoichiometry of compound is B₁₂S. Here, for the sake of comparison, one S atom was sheared between two *6c* sites, thus, giving occupation of 50%.

** the stoichiometry of the compound is B₁₂Se_{2-x}B_x.

Table II. The atomic structures and bond lengths of B₁₂S and B₁₂Se.

B ₁₂ S	Atom label (Wyckoff)	<i>x</i>	<i>y</i>	<i>z</i>	<i>B</i> _{iso} , Å ²	Site occupancy
	S1 (<i>6c</i>)	0.0000	0.0000	0.0934(2)	3.1(1)	0.55
	B1 (<i>18h</i>)	0.4419(2)	0.5581(2)	0.0437(2)	3.6(1)	1.0 ^f
	B2 (<i>18h</i>)	0.5004(3)	0.4996(3)	0.1948(2)	3.9(1)	1.0 ^f
	S-S, Å		2.230(20)			
	B2-S, Å		1.790(10)			
	B1-B2, Å		1.862(14)			
	B1-B1, Å		1.890(30)			
	B2-B2, Å		1.776(9)			
B ₁₂ Se	Atom label (Wyckoff)	<i>x</i>	<i>y</i>	<i>z</i>	<i>B</i> _{iso} , Å ²	Site occupancy
	Se1 (<i>6c</i>)	0.0000	0.0000	0.08501(1)	1.5(1)	0.52
	B1 (<i>18h</i>)	0.4360(2)	0.5640(2)	0.0481(3)	2.0(1)	1.0 ^f
	B2 (<i>18h</i>)	0.4955(3)	0.5045(3)	0.1953(3)	1.5(0)	1.0 ^f
	Se-Se, Å		2.020(30)			
	B2-Se, Å		1.872(12)			
	B1-B2, Å		1.812(14)			
	B1-B1, Å		1.830(30)			
	B2-B2, Å		1.802(10)			

^f the atom site occupancies were fixed to 1.0

Table III. The frequencies of the experimentally observed Raman bands at 633 nm excitation wavelength (ω_0) and Raman active modes predicted by CRYSTAL17 (ω_i^C), respectively. The overlapped bands groups (or too large bands) observed in experimental Raman spectra are presented by the corresponding frequency regions.

B₁₂S			B₁₂Se		
ω_0, cm^{-1}	$\omega_i^C, \text{cm}^{-1}$	modes	ω_0, cm^{-1}	$\omega_i^C, \text{cm}^{-1}$	modes
319.2	323.3	A1g	—	199.6	A1g
355.3	355.2	Eg		231.5	Eg
458.2	486.2	Eg	449.6	477.9	Eg
625.0	677.2	A1g	645.3	644.6	A1g
640-770	731.8	Eg	660-730	697.5	Eg
	754.1	Eg		711.7	Eg
	761.9	A1g	741.1	745.7	A1g
813.2	824.8	Eg	812.1	804.3	Eg
944-1170	1039.2	A1g	900-1135	1020.7	Eg
	1047.6	Eg		1024.1	A1g
	1093.7	A1g		1038.4	A1g
	1095.1	Eg		1062.8	Eg

Table IV. The lattice parameters of the optimized unit cell and atomic coordinates of B₁₂S and B₁₂Se.

B ₁₂ S	Atom label (Wyckoff)	<i>x</i>	<i>y</i>	<i>z</i>	Site occupancy
	S1 (<i>6c</i>)	0.0000	0.0000	0.1186	1.0 ^f
	B1 (<i>18h</i>)	0.4366	-0.4366	0.0479	1.0 ^f
	B2 (<i>18h</i>)	-0.1787	0.1787	0.1400	1.0 ^f
	<i>a</i> , Å	5.8966			
	<i>c</i> , Å	12.1135			
	<i>V</i> , Å ³	364.76			
B ₁₂ Se	Atom label (Wyckoff)	<i>x</i>	<i>y</i>	<i>z</i>	Site occupancy
	Se1 (<i>6c</i>)	0.0000	0.0000	0.1217	1.0 ^f
	B1 (<i>18h</i>)	0.4352	-0.4352	0.0467	1.0 ^f
	B2 (<i>18h</i>)	-0.1836	0.1836	0.1397	1.0 ^f
	<i>a</i> , Å	6.0496			
	<i>c</i> , Å	12.1603			
	<i>V</i> , Å ³	385.42			

^f all calculations were performed for site occupancies fixed to 1.0

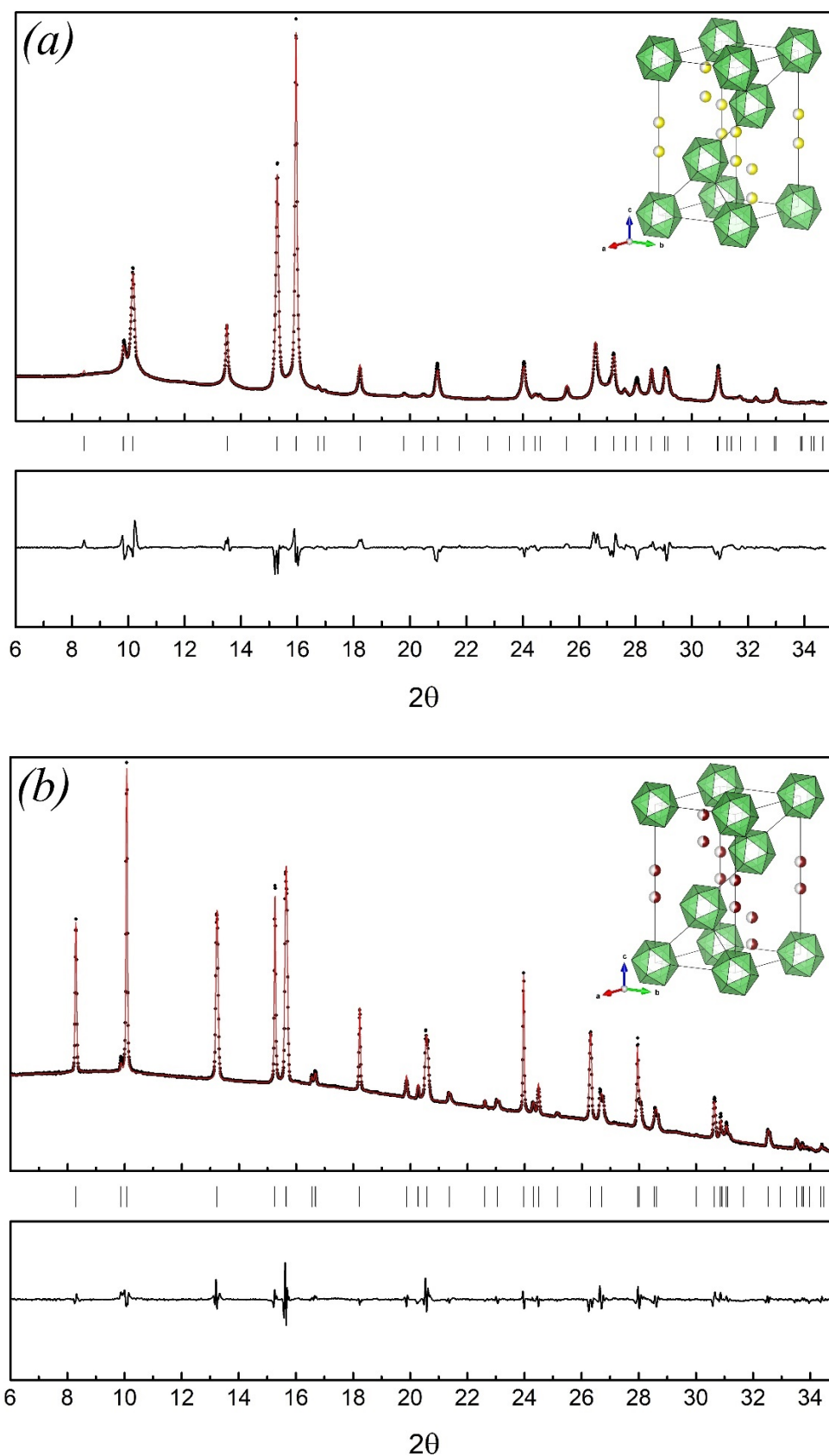


Figure 1 Rietveld full profile refinement fits of powder X-ray diffraction patterns of B₁₂S (a) and B₁₂Se (b). The insets present the unit cells of both compounds in hexagonal setting.

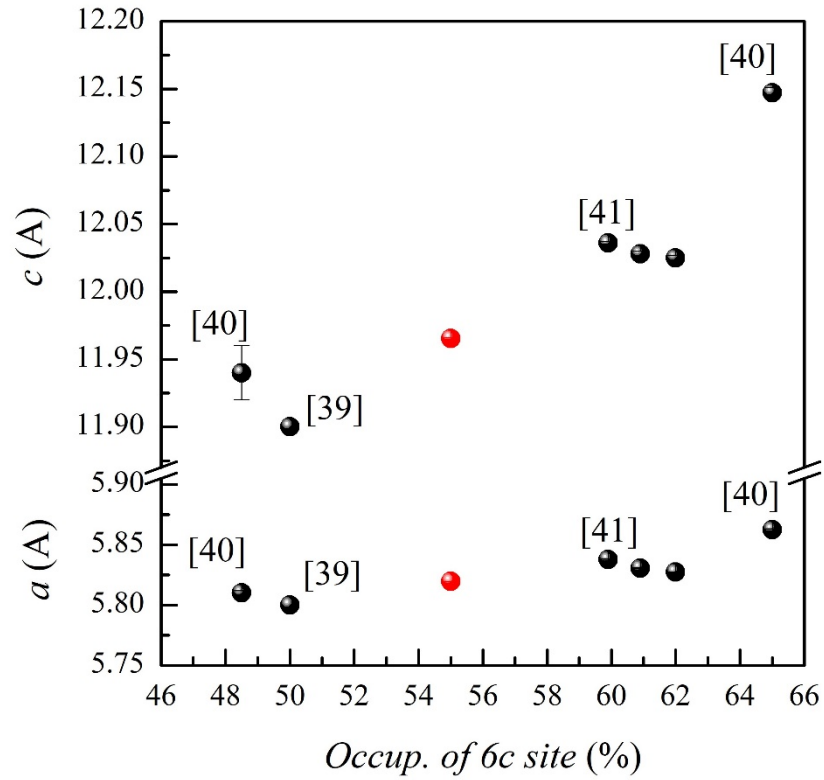


Figure 2 Comparison of B₁₂S lattice parameters *versus* occupation of 6c site by S atom obtained in present work (red balls) with literature data (black balls).

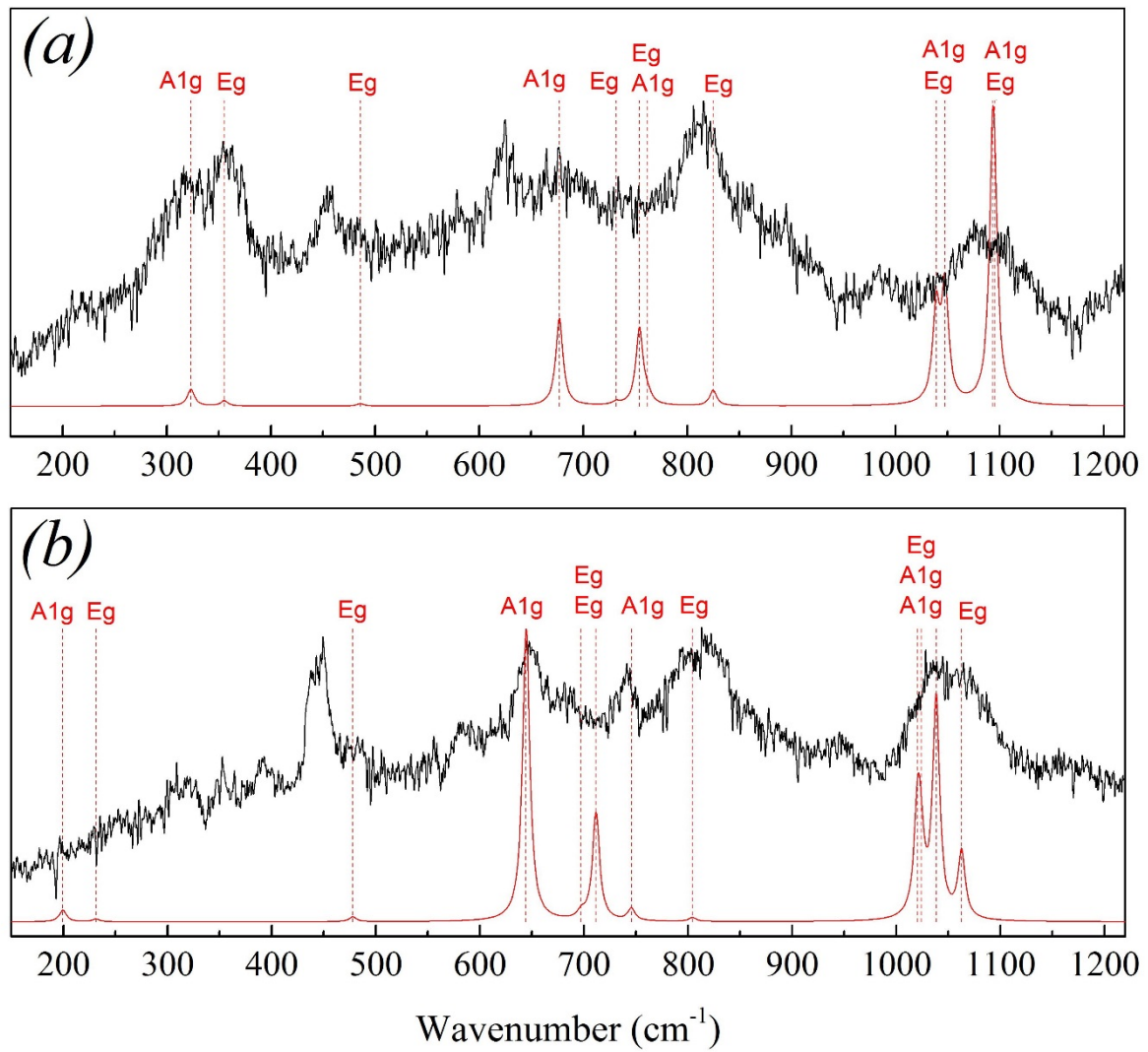


Figure 3 Experimentally observed (solid lines) and LCAO calculated (red lines) Raman bands of B₁₂S (a) and B₁₂Se (b) at ambient conditions. The positions of the predicted Raman peaks are traced by red dashed lines, the corresponding phonon modes are indicated.

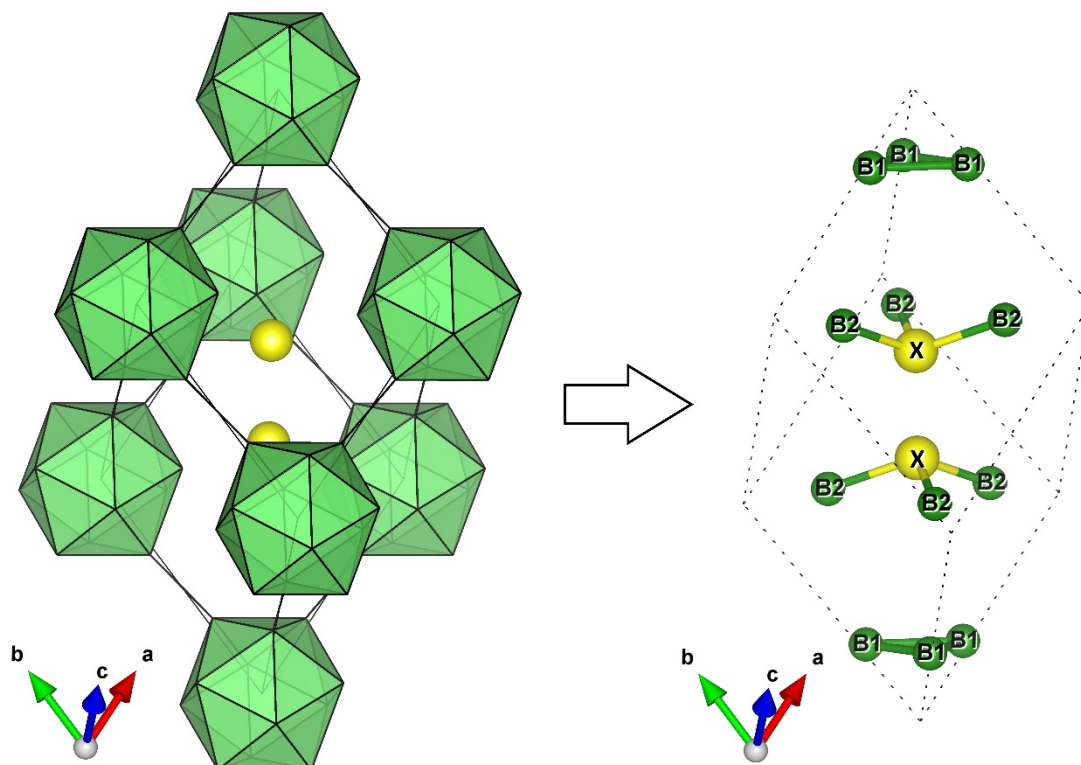


Figure 4 B₁₂X (X = S, Se) unit cell in rhombohedral setting. The polar (B1) and equatorial (B2) boron atoms of B₁₂-icosahedra are shown by green balls, S/Se atoms are shown by yellow balls.